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2. Gas particles are in random, constant, straight line motion. 3. When particles collide with each other (or any container), the collisions are said to be elastic. 4. The volume that gas particles take up is negligible. And from the distance between particles is relatively great.

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Calculate the new temperature of a gas when 1500 mL at 25°C is suddenly compressed to 500 mL. Charles's Law $K \text{ mL} / K \text{ mL} = V_1 / V_2 = T_1 / T_2$
 $100 (1500) / (298) = (500) / 1$
1 1 2 2 12. A flask contains 34.6 kPa of CO₂

Student Review Packet Answer Key

Behavior Of Gases Review 2 SECTION 2 BEHAVIOR OF GASES 1. a measure of how fast the particles of an object are moving 2. when it is heated 3. Temperature of gas particles Energy of gas particles Volume of gas particles 1) 20°C Particles have the smallest amount of energy. Volume is smallest. 2) 50°C Particles have more

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and the van der Waals equation. $(P + n^2a/V^2)(V - nb) = nRT$.
has been introduced to deal with non-ideal behavior of gases
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You usually cannot feel it, but air has pressure. The gases in Earth's atmosphere exert pressure against everything they contact. The atmosphere rises high above Earth's surface. It contains a huge number of individual gas particles. As a result, the pressure of the tower of air above a given spot on Earth's surface is substantial.

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