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1- every sample has exact same characteristic physical and chemical properties used to identify substance
2- every sample of a pure

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substance has exact same composition with the same makeup (i.e.: pure water is always 11.2% H and 88.8% oxygen by mass)

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PRESS 2011 molar mass of $I_2 = 253.80 \text{ g mol}^{-1}$
mass of $I_2 = 1.25 \times 10^{-3} \times 253.80$, i.e.

0.317 g 16 a molar mass of $PbI_2 = 460.99$

$0.1270 \text{ moles of } PbI_2 = 460.99 = 2.755 \times 10^{-4}$

mol b $Pb(NO_3)_2(aq) + MI_2(aq) \rightarrow PbI_2(s) + M(NO_3)_2(aq)$

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