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Solution Chemistry Notes. 2. Solutions • Basic definitions – Solution: homogenous mixture – Solute: substance that is dissolved – Solvent: substance that contains the solute in a solution • Process of Dissolving: – Molecules of the solvent “ pull apart ” the solute molecules, and distributes them within the solvent • How to tell the solute from the solvent – The substance that is in the greater amount is the solvent – The substance that keeps its phase is the solvent.

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Revision Notes on Solution: A solution is a homogeneous mixture of two (or more) substances, the composition of which may vary between certain limits. A solution consisting of two components is called binary solution. The component which is present in large quantity is called solvent and the component which is small in quantity is called solute.

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1. A solution is a homogeneous mixture of two or 9. more chemically non-reacting substances. The components of a solution generally

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cannot be separated by filtration, settling or centrifuging. 2. A solution may be classified as solid, liquid or a gaseous solution. 3.

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The solution which (I) Obey Raoult's law at all temperatures and concentrations (II) $H_{mix} = 0$ i.e. no heat is evolved or absorbed when components are mixed to form the solution (III) $V_{mix} = 0$ i.e. no change in volume. In ideal solution the A – B intermolecular interactions are the same as A – A and B – B inter-molecular interactions.

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Solution is a homogeneous mixture of two or more substances in same or different physical phases. The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution. Solute and Solvent In a binary solution, solvent is the component which is present in large quantity while the other

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A solution is a homogeneous mixture of two or more components in which the particle size is smaller than 1 nm. Common examples of solutions are the sugar in water and salt in water solutions, soda water, etc. In a solution, all the components appear as a single phase. There is particle homogeneity i.e. particles are evenly distributed.

Solution - Definition, Properties, Types, Videos & Examples

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01. Solutions - PhysicsWallah

Redox reactions, conductance in electrolytic solutions, specific and molar conductivity, variations of conductivity with concentration, Kohlrausch's Law, electrolysis and law of electrolysis (elementary idea), dry cell-electrolytic cells and Galvanic cells, lead accumulator, EMF of a cell, standard electrode potential, Nernst equation and its application to chemical cells, Relation between Gibbs energy change and emf of a cell, fuel cells, corrosion.

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basis of number of components a solution of two components is called binary solution.

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Kerala Plus Two Chemistry Notes Chapter 2 Solutions. Solution A homogeneous mixture of two or more sub-stances is knows as solution. Solute : The substance present in smaller amount in a solution is called solute. Solvent: The substance present in larger amount in a solution is called solvent. Mass Percentage: It is the amount of solute in grams dissovded per 100 g of solution.

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Chemistry Solution Notes CBSE Class 12 pdf (Important ...
Solutions Class 12 Notes Chapter 2 A solution comprises a solute and

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a solvent. It is defined as a homogeneous mixture of two or more than two substances. They can be classified into three types:

CBSE Class 12 Chemistry Notes Chapter 2 Solutions

When a metal conducts it is the electrons that are moving through the metal. When a salt solution conducts it is the ions in the solution that move towards the electrodes carrying the electrons. During electrolysis oxidation of the non-metal ions always occurs at the anode and reduction of the metal or hydrogen ions occurs at the cathode.

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