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the amounts of reactants and products that take part in a chemical reaction.

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chemical reaction. to begin, check that this

equation is shown: $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} +$

3CO_2 1. look at the coefficients (such as the

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answers equation: $\text{Fe}_2\text{O}_3(\text{s}) + 3\text{CO}(\text{g}) \rightarrow 2\text{Fe}(\text{s})$

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Quiz Answers equation: $\text{Fe}_2\text{O}_3(\text{s}) + 3$

$\text{CO}(\text{g}) \rightarrow 2 \text{Fe}(\text{s}) + 3\text{CO}_2(\text{g})$ There is a

Question to solve under the equation.

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Stoichiometry Gizmo™ allows you to try

your hand at figuring out the amounts of

reactants and products that take part in a

chemical reaction.

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The Stoichiometry Gizmo™ allows you to

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of reactants and products that take part in

a chemical reaction. To begin, check that

this equation is shown: $\text{Fe}_2\text{O}_3 + 3\text{CO}$

$\rightarrow 2\text{Fe} + 3\text{CO}_2$ 1. Look at the coefficients

(such as the “ 3 ” in 3CO) in front of

each substance in the equation.

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converted into units of the answer. Tiles
can be flipped, and answers can be
calculated once the appropriate unit
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Since 20 moles NO_2 are produced, and it takes 3 moles O_2 to produce 4 moles of NO_2 , the answer would be 15 moles. $15 \text{ moles} = 20 \text{ moles } \text{NO}_2 \times (3 \text{ moles } \text{O}_2 / 4 \text{ moles } \text{NO}_2)$ --- This is the molar ratio between the two.

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